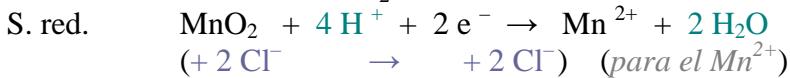
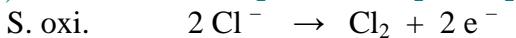
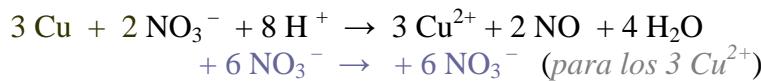
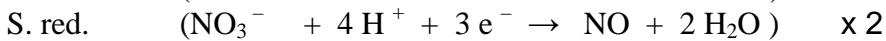
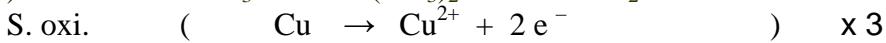
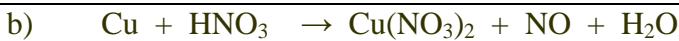


Resueltos (medio ácido)

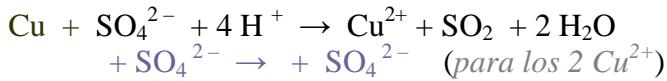
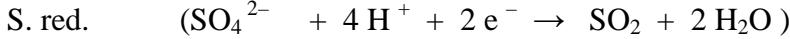
1.- Ajusta las siguientes reacciones por el método del ión-electrón en medio ácido
Para ajustar el oxígeno, se añade el H₂O necesaria en el lado que falta O, y se ajusta con H⁺ en el otro lado.



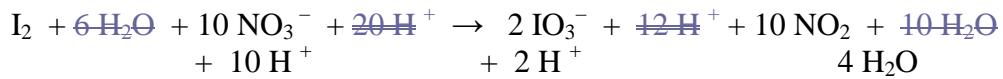
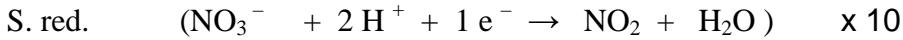
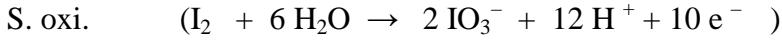
Ajustada: $4 \text{HCl} + \text{MnO}_2 \rightarrow \text{MnCl}_2 + \text{Cl}_2 + 2 \text{H}_2\text{O}$



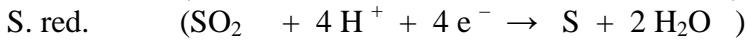
Ajustada: $3 \text{Cu} + 8 \text{HNO}_3 \rightarrow 3 \text{Cu}(\text{NO}_3)_2 + 2 \text{NO} + 4 \text{H}_2\text{O}$



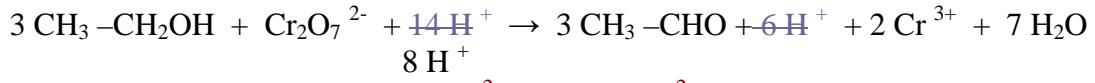
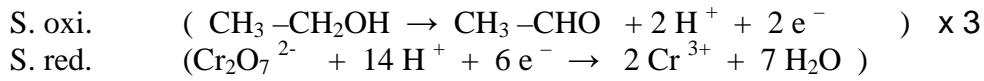
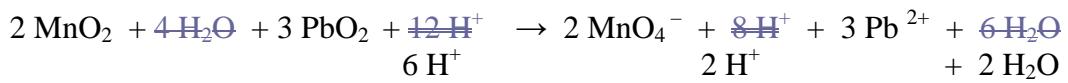
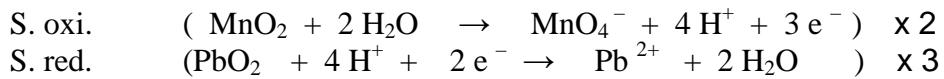
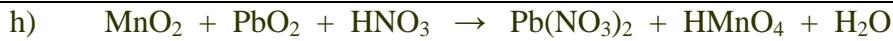
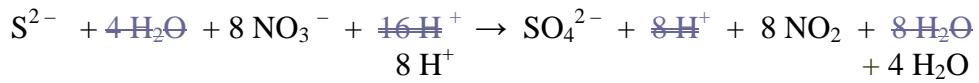
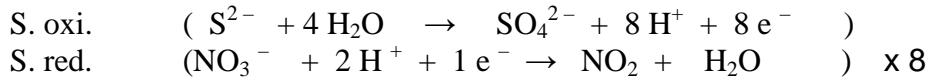
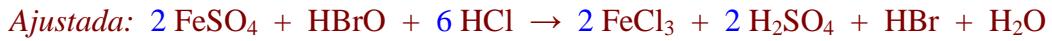
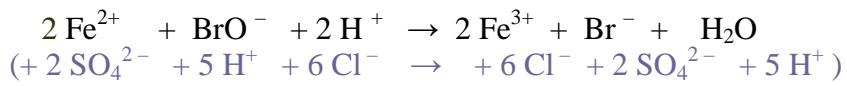
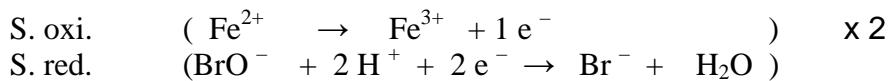
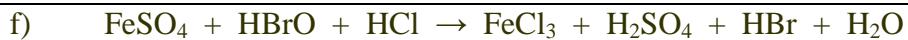
Ajustada: $\text{Cu} + 2 \text{H}_2\text{SO}_4 \rightarrow \text{CuSO}_4 + \text{SO}_2 + 2 \text{H}_2\text{O}$



Ajustada: $\text{I}_2 + 10 \text{HNO}_3 \rightarrow 2 \text{HIO}_3 + 10 \text{NO}_2 + 4 \text{H}_2\text{O}$



Ajustada: $2 \text{H}_2\text{S} + \text{SO}_2 \rightarrow 3 \text{S} + 2 \text{H}_2\text{O}$



Ajustada:



(Se oxidan tanto el Cu^+ a Cu^{2+} , como el S^{2-} a S^{4+} . Se reduce el N^{5+} a N^{2+})

